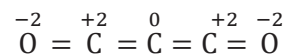


THE PARADOX OF FRACTIONAL OXIDATION

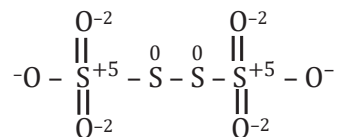
Carbon Suboxide



The average oxidation number of carbon atoms:

$$\frac{(2+2+0)}{3} = \frac{4}{3}$$

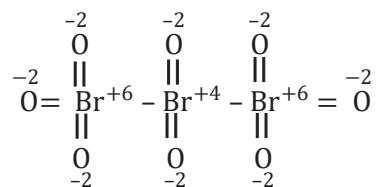
Tetrathionate Ion



The average oxidation number of sulphur atom in the ion:

$$\frac{5 + 5 + 0 + 0}{4} = 2.5$$

Tribromooxaoxide



The average oxidation number of bromine atom:

$$\frac{(6 + 6 + 4)}{3} = \frac{16}{3}$$