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Question 1:

(i) Write the electron-dot structures for sodium, oxygen and magnesium.

- (ii) Show the formation of Na₂O and MgO by the transfer of electrons.
- (iii) What are the ions present in these compounds?

Answer 1:

(i) Electron – dot structure for Sodium:



Electron – dot structure for Oxygen:



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Electron – dot structure for Magnesium:



(ii). Formation of Na₂O by transfer of electron:



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Formation of Na₂O by transfer of electron:



(iii). Ions present in these compounds are Mg^{2+} , O^{2-} and Na^+ .

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Question 2:

Why do ionic compounds have high melting points?

Answer 2:

Ionic compounds have high melting and boiling points. Because ionic compounds are formed by the attraction force of two opposite ions and a considerable amount of energy is required to break this strong inter-ionic attraction.

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