Mendeleev's periodic table is based on:-#Atomic number#Increasing order of number of protons#Electronic configuration#None of the above
Which of the following is/are Dobereiners triad:-
(a) P, As, Sb
 (b) Cu, Ag, Au
(c) Fe, Co,
 (d) S, Se, Te
Correct answer is :- #a and b#b and c#a and d#All
Which of the following sets of elements follows Newland's octave rule :-#Be, Mg, Ca#Na, K,Rb#F, Cl, Br#B, AL, Ga
Which are correct match:-
(a) Eka silicon - Be
 (b) Eka aluminium - Ga
(c) Eka manganese - Tc
 (d) Eka scandium - B#b, c#a, b, d#a, d#All
Atomic wt. of P is 31 and Sb is 120. What will be the atomic wt. of as, as per Dobereiners triad rule:-#151#75.5#89.5#Unpredictable
The places that were left empty by Mendeleev's were, for:-#Aluminium & Silicon#Galium and germinium#Arsenic and antimony#Molybdenum and tungstun
Which is not anomalous pair of elements in the Mendeleev's periodic table:-#Ar and K#Co and Ni#Te and I#Al and Si
The law of triads is applicable to :-#Os, Ir, Pt#Ca, Sr, Ba#Fe, Co, Ni#Ru, Rh, Pt
Elements which occupied position in the other meyer curve, on the peaks, were:-#Alkali metals#Highly electro positive elements#Elements having large atomic volume#All
In a period the elements are arranged in :#Decreasing order of nuclear charge#Decreasing order of No. of electrons#Increasing order of nuclear charge#In order of same nuclear charge
Which of the following statement is wrong :#No inert gas is present in 7^{th} period#3^{rd} periodcontains 18 elements#1^{st} period contains two non metals#In p-block, metal, nonmetal and metalloids are present
Which of the following element was absent in the Mendeleev's periodic table:- #Tc #Si#B#F
IUPAC name of the element placed just after actinide series :-#Unniltrium#Unnilpentilium#Unnilquadium#Ununbium
Which statement is wrong for the long form of periodic table:- #Number uf periods are 7 and groups 18#No. of valence shell electrons in a period are, same#III B group contains 32 elements#Lanthanides and actinides are placed in same group
The elements which.· are cited as an example to proove the validity of mendeleev's periodic low are#H, He#Ga, Sc#Co, Ni#Zr, Hf
Which pair of successive elements follows increasing order of at6mic weight in mendeleev's periodic table. #Argon and potassium#Lithium and Beryilium #Cobalt and nickel#Tellurium and iodine
Which of the following statement is false :- #Elements of ns^2np^6 electronic configuration lies in 1^{st} to 6^{th} period #Typical elements lies in 3^{rd} period#The seventh period will accommodate thirty two elements #Boron and silicon are diagonally related
Among the Lanthanides the one obtained by synthetic method is :-#Lu#Pm#Pr#Ce

Which of the following set of elements belongs to same period :#Zn, Cd, Hg#Fr, Ra, U#K, Ca, Ag#None
The element with atomic number Z = 115 will be placed in:-#7^{th} period, IA group#8^{th} period, NA group#7^{th} period, VA group#6^{th} period, VB group

Elements upto atomic no. 112 have been discovered till now. What will be the electronic configuration of the element possessing atomic no 108 :-
[Rn]$5f^{14}$ $6d^6$ $7s^2$
$6f^{14}$ $7d^8$ 7^2
[Rn] $5f^{14}$ $6d^8$ $7s^0$
[Xe] $4f^{14}$ $5d^8$ $6s^2$

In 6^{th} period of the modern periodic table, electronic energy levels are in the order #6s, 4f, 5d, 6p#6s, 6p, 4f, 5d#4f, 5d, 6s, 6p#None
Out of first 100 elements no. of elements having electrons in 3d orbital (in their complete electronic configuration) are :-#80#100#40#60
The IUPAC name of the· element which is placed after Db_{105} is the periodic table, will be :-#Un nil pentium#Un un nilium#Un nil hexium#Un nil quadium
The element with the electronic configuration $ns^2(n-1)s^2p^6d^0(n-2)s^2p^6d$ $^{10}f^7$ 1ies in the :-#s- block#p - block#d - block#f – block
The element with atomic number Z= 118 will be :#Noble gas#Transition metal#Alkali metal#Alkaline earth metal
The atom having the valence shell electronic configuration 4s² 4p² would be in:-#Group II A and period 3#Group II B and period 4#Group N A and period 4#Group N A and period 3

The electronic configuration of d-block elements is exhibited by :-
$ns^{1-2}(n-1)d^{1-10}$
ns^2 $(n-1)$ d^{10}
$(n-1)d^{10}s^2$
ns^2np^5
The electronic configuration of the element with atomic number 109 if discovered will be:-
$(n-1)d^7ns^2$
$(n-1)d^9ns^2$
nd^7ns^2
$(n-1)d^5ns^2np^2$

The element having electronic configuration $4f^{14}5d^06s^2$ belongs to :-

#d-block, 12^{th} group#f-block, III B group#f-block, 14^{th} group#s-block, 2^{nd} group

Element with the electronic configuration given below, belong to which group in the periodic table $1s^2$, $2s^22p^6$, $3s^23p^63d^{10}$, $4s^24p^64d^{10}$, $5s^25p^3$

#3rd#5th#15th#17th

$4d^35s^2$ configuration belongs to which group :- .

#II A#II B#VB#III B

Which of the. following electronic configuration belongs to inert gas elements :-
#ns^2 $(n-1)d^{10}$
#ns^2 $(n-1)s^2p^6$
#ns^2np^6
#None

From atomic number 58 to 71, elements are placed in:-#5th period and III A group#6th period and III B group#Separate period and group#7th period and. N B group
True statement is :-#All the transuranic elements are synthetic elements#Elements of third group are called bridge elements#Element of $1s^2$ configuration is placed in II A group#Electronic configuration of elements of a group is same
Elements having ns^2 np^6 valence shell electronic configuration lies in :-#'0' gp. and 1st–7th period#18th gp. and 2nd–6th period#18th gp. and 1st–6th period#All are correct
Which of the following match is correct:-#Last natural element - Uub#General electronic configuration of IA group – ns²#Inert gas elements lies in 2nd – 6th period#Typical elements- 3rd period elements

The electronic configuration of elements X and Z are $1s^2$ $2s^2$ $2p^6$ $3s^2$ $3p^5$ and $1s^2$ $2s^2$ $2p^5$ respectively. What is the position of element X with respect to position of Z in the periodic table -
#Just below Z#Just above Z#Left to the Z#right to the Z

Which of the following sequence contains atomic number of only representative elements#55, 12, 18, 53#13, 33, 54, 83#3, 33, 53, 87#22, 33, 55, 66
Uranium (At No. - 92) is the last natural element in the periodic table. The last element of the periodic table which is recently discovered is Uub. What will be the total number of transuranic elements in the periodic table :-#21#20#11#12
Which-two elements are in same period as well as same group of modem periodic table :-#Z = 23, Z = 31#Z = 65, Z = 66#Z = 52, Z = 87#Z = 58, Z = 46

Which of the following statement is not correct for given electronic configuration $1s^2$, $2s^22p^6$, $3s^23p^63d^{10}$, $4s^24p^64d^{10}4f^{14}$, $5s^25p^65d^{10}$, $6s^2$
#It belongs to IIB group and 6^{th} period#It is liquid at room temperature#It is a transition element#It is not used in high temperature thermometer

General electronic configuration of outermost
and penultimate shell is $(n-1)\ s^2\ (n-1)p^6(n-1)$
$d^x\ ns^2$. If n = 4 and x = 5, then number of protons
in the nucleus will be :-
#> 25#< 24#25#30

An ion M⁺³ has electronic configuration [Ar]3d¹⁰
4s² element M belongs to :-#s-block#p-block#d-block#f-block

What is the atomic number of element having maximum number of unpaired e^- in 4p subshell:-#33#17#53#15

The formula for effective nuclear charge is (if is screening constant)
#Z – S#Z + S#Z S⁻¹#Z S

According to Slater rule, Effective nuclear charge in group generally :- #Increases down the group#Decreases down the group#Remains constant#First increases than decreases
In sodium atom the screening is due to :-#3s²,
3p⁶#2s¹#1s², 2s²,
2p⁶ #1s², 2s²

If the difference in atomic size of :
Na – Li = x; Rb – K = y; Fr – Cs = z
Then correct order will be:-#X = y = Z#X > y > Z#X< y < Z#X< y << Z
The correct order of size would be:-#Ni < Pd > Pt#Pd < Pt < Ni#Pt > Ni > Pd#Pd > Pt > Ni

Which of the following order of radii is correct#Li < Be < Mg#H⁺ < Li⁺ < H⁻
#O < F < Ne#Na⁺ > F⁻ > O⁻²

K⁺, Ar, Ca²⁺ and S²⁻contains -#Same electronic configuration and atomic volume#Different electronic configuration but same IP.#Same electronic configuration but different atomic volume#None

Answers

1.(4)**2.**(3)**3.**(1)**4.**(1)**5.**(2)**6.**(2)**7.**(4)
8.(2)**9.**(4)**10.**(3)**11.**(2)**12.**(1)**13.**(3)**14.**(2)
15.(2)**16.**(2)**17.**(1)**18.**(2)**19.**(2)**20.**(3)**21.**(1)
22.(1)**23.**(1)**24.**(3)**25.**(4)**26.**(1)**27.**(3)**28.**(1)
29.(1)**30.**(2)**31.**(3)**32.**(3)**33.**(3)**34.**(2)**35.**(1)
36.(2)**37.**(4)**38.**(1)**39.**(3)**40.**(2)**41.**(2)**42.**(3)
43.(3)**44.**(2)**45.**(1)**46.**(1)**47.**(3)**48.**(3)**49.**(2)
50.(1)**51.**(2)**52.**(3)**53.**(4)