Mendeleev's periodic table is based on:-#Atomic number#Increasing order of number of protons#Electronic configuration#None of the above

Which of the following is/are Dobereiners triad:-<br />(a) P, As, Sb<br /> (b) Cu, Ag, Au<br />(c) Fe, Co, <br /> (d) S, Se, Te<br />Correct answer is :- #a and b#b and c#a and d#All

Which of the following sets of elements follows Newland's octave rule :-#Be, Mg, Ca#Na, K,Rb#F, Cl, Br#B, AL, Ga

Which are correct match:- <br />(a) Eka silicon - Be<br /> (b) Eka aluminium - Ga<br />(c) Eka manganese - Tc<br /> (d) Eka scandium - B#b, c#a, b, d#a, d#All

Atomic wt. of P is 31 and Sb is 120. What will be the atomic wt. of as, as per Dobereiners triad rule:-#151#75.5#89.5#Unpredictable

The places that were left empty by Mendeleev's were, for:-#Aluminium & Silicon#Galium and germinium#Arsenic and antimony#Molybdenum and tungstun

Which is not anomalous pair of elements in the Mendeleev's periodic table:-#Ar and K#Co and Ni#Te and I#Al and Si

The law of triads is applicable to :-#Os, Ir, Pt#Ca, Sr, Ba#Fe, Co, Ni#Ru, Rh, Pt

Elements which occupied position in the other meyer curve, on the peaks, were:-#Alkali metals#Highly electro positive elements#Elements having large atomic volume#All

In a period the elements are arranged in :#Decreasing order of nuclear charge#Decreasing order of No. of electrons#Increasing order of nuclear charge#In order of same nuclear charge

Which of the following statement is wrong :#No inert gas is present in 7th period#3rd periodcontains 18 elements#1st period contains two non metals#In p-block, metal, nonmetal and metalloids are present

Which of the following element was absent in the Mendeleev's periodic table:- #Tc #Si#B#F

IUPAC name of the element placed just after actinide series :-#Unniltrium#Unnilpentilium#Unnilquadium#Ununbium

Which statement is wrong for the long form of periodic table:- #Number uf periods are 7 and groups 18#No. of valence shell electrons in a period are, same#III B group contains 32 elements#Lanthanides and actinides are placed in same group

The elements which.· are cited as an example to proove the validity of mendeleev's periodic low are#H, He#Ga, Sc#Co, Ni#Zr, Hf

Which pair of successive elements follows increasing order of at6mic weight in mendeleev's periodic table. #Argon and potassium#Lithium and Beryilium

#Cobalt and nickel#Tellurium and iodine

Which of the following statement is false :- #Elements of ns2np6 electronic configuration lies in 1st to 6th period #Typical elements lies in 3rd period#The seventh period will accommodate thirty two elements #Boron and silicon are diagonally related

Among the Lanthanides the one obtained by synthetic method is :-#Lu#Pm#Pr#Ce

Which of the following set of elements belongs to same period :#Zn, Cd, Hg#Fr, Ra, U#K, Ca, Ag#None

The element with atomic number Z = 115 will be placed in:-#7th period, IA group#8th period, NA group#7th period, VA group#6th period, VB group

Elements upto atomic no. 112 have been discovered till now. What will be the electronic configuration of the element possessing atomic no 108 :-

# [Rn]5fl4 6d6 7s2

# 6f14 7d8 72

# [Rn] 5f14 6d8 7s0

# [Xe] 4f14 5d8 6s2

In 6th period of the modern periodic table, electronic energy levels are in the order #6s, 4f, 5d, 6p#6s, 6p, 4f, 5d#4f, 5d, 6s, 6p#None

Out of first 100 elements no. of elements having electrons in 3d orbital (in their complete electronic configuration) are :-#80#100#40#60

The IUPAC name of the· element which is placed after Db105 is the periodic table, will be :-#Un nil pentium#Un un nilium#Un nil hexium#Un nil quadium

The element with the electronic configuration ns2(n–1)s2p6d0(n–2)s2p6d 10f7 1ies in the :-#s- block#p - block#d - block#f – block

The element with atomic number Z= 118 will be :#Noble gas#Transition metal#Alkali metal#Alkaline earth metal

The atom having the valence shell electronic configuration 4s<sup>2</Sup> 4p<sup>2</Sup> would be in:-#Group II A and period 3#Group II B and period 4#Group N A and period 4#Group N A and period 3

The electronic configuration of d-block elements is exhibited by :-

# ns1–2(n–1)d1–10

# ns2 (n–1) d10

# (n–1)d10s2

# ns2np5

The electronic configuration of the element with atomic number 109 if discovered will be:-

# (n–1)d7ns2

# (n–1)d9ns2

# nd7ns2

# (n–1)d5ns2np2

The element having electronic configuration 4f145d06s2 belongs to :-

#d-block, 12th group#f-block, III B group#f-block, 14th group#s-block, 2nd group

Element with the electronic configuration given below,

belong to which group in the periodic table 1s2, 2s22p6,

3s23p63d10, 4s24p64d10, 5s25p3

#3rd#5th#15th#17th

4d35s2 configuration belongs to which group :- .

#II A#II B#VB#III B

Which of the. following electronic configuration belongs to inert gas elements :-

#ns2 (n–1)d10

#ns2 (n–1)s2p6

#ns2np6

#None

From atomic number 58 to 71, elements are placed in:-#5th period and III A group#6th period and III B group#Separate period and group#7th period and. N B group

True statement is :-#All the transuranic elements are synthetic elements#Elements of third group are called bridge elements#Element of 1s2 configuration is placed in II A group#Electronic configuration of elements of a group is same

Elements having ns2 np6 valence shell electronic configuration lies in :-#'0' gp. and 1st–7th period#18th gp. and 2nd–6th period#18th gp. and 1st–6th period#All are correct

Which of the following match is correct:-#Last natural element - Uub#General electronic configuration of IA group – ns<sup>2</sup>#Inert gas elements lies in 2nd – 6th period#Typical elements- 3rd period elements

The electronic configuration of elements X and Z are

1s2 2s2 2p6 3s2 3p5 and 1s2 2s2 2p5 respectively. What

is the position of element X with respect to position of

Z in the periodic table –

#Just below Z#Just above Z#Left to the Z#right to the Z

Which of the following sequence contains atomic number of only representative elements#55, 12, 18, 53#13, 33, 54, 83#3, 33, 53, 87#22, 33, 55, 66

Uranium (At No. - 92) is the last natural element in the periodic table. The last element of the periodic table which is recently discovered is Uub. What will be the total number of transuranic elements in the periodic table :-#21#20#11#12

Which-two elements are in same period as well as same group of modem periodic table :-#Z = 23, Z = 31#Z = 65, Z = 66#Z = 52, Z = 87#Z = 58, Z = 46

Which of the following statement is not correct for

given electronic configuration 1s2, 2s22p6, 3s23p6

3d10, 4s24p64d104f1 4, 5s25p65d10, 6s2

#It belongs to IIB group and 6th period#It is liquid at room temperature#It is a transition element#It is not used in high temperature thermometer

General electronic configuration of outermost and penultimate shell is (n–1) s<sup>2</sup> (n–1)p<sup>6</sup> (n–1)d<sup>x</sup> ns<sup>2</sup>. If n = 4 and x = 5, then number of protons in the nucleus will be :-#> 25#< 24#25#30

An ion M<sup>+3</sup> has electronic configuration [Ar]3d<sup>10</sup> 4s<sup>2</sup> element M belongs to :-#s-block#p-block#d-block#f-block

What is the atomic number of element having maximum number of unpaired e– in 4p subshell:-#33#17#53#15

The formula for effective nuclear charge is (if  is screening constant)#Z – S#Z + S#Z S<sup>–1</sup>#Z S

According to Slater rule, Effective nuclear charge in group generally :- #Increases down the group#Decreases down the group#Remains constant#First increases than decreases

In sodium atom the screening is due to :-#3s<sup>2</sup>, 3p<sup>6</sup>#2s<sup>1</sup>#1s<sup>2</sup>, 2s<sup>2</sup>, 2p<sup>6</sup> #1s<sup>2</sup>, 2s<sup>2</sup>

If the difference in atomic size of :<br />Na – Li = x; Rb – K = y; Fr – Cs = z<br />Then correct order will be:-#X = y = Z#X > y > Z#X< y < Z#X< y << Z

The correct order of size would be:-#Ni < Pd > Pt#Pd < Pt < Ni#Pt > Ni > Pd#Pd > Pt > Ni

Which of the following order of radii is correct#Li < Be < Mg#H<sup>+</sup> < Li<sup>+</sup> < H<sup>-</sup>#O < F < Ne#Na<sup>+</sup> > F<sup>-</sup> > O<sup>-2</sup>

K<sup>+</sup>, Ar, Ca<sup>2+</sup> and S<sup>2-</sup>contains -#Same electronic configuration and atomic volume#Different electronic configuration but same IP.#Same electronic configuration but different atomic volume#None

**Answers**

**1.**(4)

**2.**(3)

**3.**(1)

**4.**(1)

**5.**(2)

**6.**(2)

**7.**(4)

**8.**(2)

**9.**(4)

**10.**(3)

**11.**(2)

**12.**(1)

**13.**(3)

**14.**(2)

**15.**(2)

**16.**(2)

**17.**(1)

**18.**(2)

**19.**(2)

**20.**(3)

**21.**(1)

**22.**(1)

**23.**(1)

**24.**(3)

**25.**(4)

**26.**(1)

**27.**(3)

**28.**(1)

**29.**(1)

**30.**(2)

**31.**(3)

**32.**(3)

**33.**(3)

**34.**(2)

**35.**(1)

**36.**(2)

**37.**(4)

**38.**(1)

**39.**(3)

**40.**(2)

**41.**(2)

**42.**(3)

**43.**(3)

**44.**(2)

**45.**(1)

**46.**(1)

**47.**(3)

**48.**(3)

**49.**(2)

**50.**(1)

**51.**(2)

**52.**(3)

**53.**(4)